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11—1 Review and Reinforcement (continued) Solve each of the following problems as directed. Show all your work. 13. How many moles Of magnesium are required to react with 2.0 mol of hydrochloric acid? The equation for this reaction is $Mg + 2HCl \rightarrow MgCl_2 + H_2$ 14. Aluminum reacts with HCl to produce aluminum chloride ($AlCl_3$) and hydrogen gas.

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11—2 Review and Reinforcement Solving Stoichiometry Problems If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line provided. 1. 2. 5. The major types of stoichiometry problems are mass-mass, mass-volume, and volume-volume.

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11. will remain constant. The equilibrium position, however, is shifted to the right, more product. 12. If the concentration Of a reactant is increased. the reaction Will shift to a'nsome some of the added reactant and product will be produced. If reactant is removed, the reaction will shift to more reactant to compensate for the loss. 13.

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11—1 Review and Reinforcement Stoichiometry Complete the following sentences by 'filling in the appropriate word or phrase from the list below. Each word or phrase maybe used once, more than once, or not at all reactants molar ratio quantitative actual particles. consdrvption of matter coefficients subscripts mass

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10—2 Review and Reinforcement (continued) 9. 0.0621 mol SS = (Q \$ gK2S 10. 65.8gMg= l. atoms Mg Solve each of the following problems as directed. Show all your work. 11. Calculate the number Of atoms in 0.40 mol of sulfur- 12. Calculate the number of atoms in 2.30 mol of silver. G.oux 13. You have a tank of NOZ gas.

10-2 Review and Reinforcements

SECTION 11-1 REVIEW AND REINFORCE The Water Cycle ^ Understanding Main ideas ... L 11. sludge 1, 12. sewage Science Explorer Earth Science I Unit 4 Resources 39 .

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